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## **Chapter 11 Stoichiometry Study Guide**

TEACHER GUIDE AND ANSWERS Study Guide - Chapter 11 - Stoichiometry

Section 11.1 What is stoichiometry? 1.

true 2. true 3. false 4. true 5. true 6. 2,

2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4,

4, 72.08 10. methanol and oxygen gas

11. carbon dioxide and water 12. 160.10

g 13. 160.10 g 14. They are equal. 15. A

mole ratio is a ratio between the numbers of moles

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### **VIBRATIONS AND WAVES**

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-Stoichiometry (Rough outline of the chapter, please use the book, notes & homework to study.) 11.1 Defining Stoichiometry Vocab • stoichiometry • mole ratio Concepts Using Balanced Equations • Number of Atoms • Number of Molecules • Number of Moles • Mass o Law of Conservation of Mass • Volume

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CHAPTER 11 SECTIONS 1 Defining Stoichiometry 2 Stoichiometric Calculations 3 Limiting Reactants 4 Percent Yield LaunchLAB What evidence can you observe that a reaction has stopped? During a chemical reaction, reactants are consumed as new products form. In this lab, you will look for signs a chemical reaction has stopped. Steps in Stoichiometric Calculations

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15.2 CHAPTER 11: STOICHIOMETRY.  
MOLE TO MOLE RATIO. When nitrogen and hydrogen gas are heated under the correct conditions, ammonia gas ( $\text{NH}_3$ ) is formed. a. RXN:  $1. \text{N}_2 + 3. \text{H}_2 \rightarrow 2. \text{NH}_3$ . b. How many moles of nitrogen react with three moles of hydrogen? 1 mol  $\text{N}_2$  3 mol  $\text{H}_2$  1 mol  $\text{N}_2$ . 3 mol  $\text{H}_2$ . c.

## **CHAPTER 11: STOICHIOMETRY**

368 Chapter 11 • Stoichiometry Section 11.11.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

## **Chapter 11: Stoichiometry**

Chapter 11 Stoichiometry.  
stoichiometry. mole ratio. excess reactant. limiting reactant. The study of

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quantitative relationships between the amounts of.... In a balanced equation, the ratio between the numbers of moles.... A reactant that remains after a chemical reaction stops.

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms

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of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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reactant. limiting reactant. The study of  
quantitative relationships between the  
amounts of.... In a balanced equation,  
the ratio between the numbers of  
moles.... A reactant that remains after a  
chemical reaction stops.

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